Atomic structure worksheet 8th grade answer key 2020 list pdf download

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CHEMICAL ANALYSIS - 3

Use the test results given below to identify the two ions which are present in the compounds and which are responsible for the results given.

1. Compound A gives a yellow colour in a flame test. A solution of the compound forms a cream precipitate when dilute nitric acid is added followed by silver nitrate solution.

   Ion 1 ________________ Ion 2 ________________ (2)

2. Compound B dissolves in dilute nitric acid producing a gas which turns lime water milky. A solution of compound B forms a green precipitate when sodium hydroxide solution is added.

   Ion 1 ________________ Ion 2 ________________ (2)

3. When compound C is warmed with sodium hydroxide solution it forms a gas which turns damp litmus paper blue. A solution of the compound forms a white precipitate when dilute hydrochloric acid is added followed by barium solution.

   Ion 1 ________________ Ion 2 ________________ (2)

4. A solution of compound D forms a blue precipitate when sodium hydroxide solution is added. A white precipitate is formed when dilute nitric acid followed by silver nitrate solution is added to a solution of compound D.

   Ion 1 ________________ Ion 2 ________________ (2)

5. Compound E gives a lilac colour when used in a flame test. When a solution of compound E is heated with sodium hydroxide solution and aluminium powder a gas is formed which turns damp red litmus paper blue.

   Ion 1 ________________ Ion 2 ________________ (2)

6. A solution of compound F turns blue litmus red. The solution of compound F forms a pale yellow precipitate when dilute nitric acid is added followed by silver nitrate solution.

   Ion 1 ________________ Ion 2 ________________ (2)

7. Compound G gives a brick-red (orangey-red) colour in a flame test. A solution of compound G turns litmus paper blue.

   Ion 1 ________________ Ion 2 ________________ (2)

Part C: Electron Configuration

12. How many electrons can each level hold? 1st = ____ 2nd = ____ 3rd = ____

13. What term is used for the electrons in the outermost shell or energy level? ____________

14. Scientists use two types of diagrams to show the electron configuration for atoms. Follow your teacher's directions to complete the diagrams.

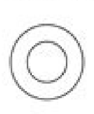
Sulfur
Atomic # = 16
Atomic Mass = 32
Protons = ___
Neutrons = ___
Electron = ___

Bohr Diagram
Shows all electrons

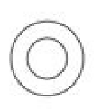
Lewis Structure
Shows valence electrons

S

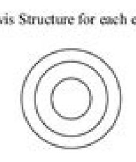
15. Calculate the missing information and then draw the Bohr Diagram and Lewis Structure for each element.



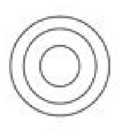
Li
Atomic # = 3
Mass # = 7
# of P = ___
# of N = ___
# of E = ___



Ne
Atomic # = 10
Mass # = 20
# of P = ___
# of N = ___
# of E = ___

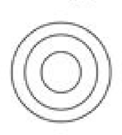


Mg
Atomic # = 12
Mass # = 24
# of P = ___
# of N = ___
# of E = ___



Cl
Atomic # = 17
Mass # = 35
# of P = ___
# of N = ___
# of E = ___

He
Atomic # = 2
Mass # = 4
# of P = ___
# of N = ___
# of E = ___



Si
Atomic # = 14
Mass # = 28
# of P = ___
# of N = ___
# of E = ___

16. Answer the questions below based on the elements in question #15.

(1) Which elements had a filled outermost shell? ____ ____

(2) Which element would be most likely to lose electrons in a chemical bond? ____

(3) Which element would be most likely to gain electrons in a chemical bond? ____

(4) Which elements are not likely to bond with other elements? ____ ____ Why? ____________

Name ______________ Date ______________ Class ______________

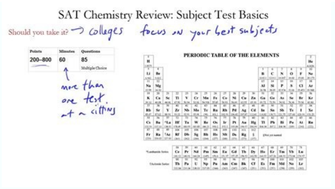
4.1 DEFINING THE ATOM

Section Review

Objectives

Vocabulary

Part A Completion



Exploring Atoms Name _____

Watch the "Atoms" movie at http://www.brainpop.com/ to complete this section.

1. List three items that are made of atoms. _____

2. The _____ _____ of _____ shows and predicts every element in the universe.

3. _____ argued that everything in the world was made up of particles so small they could not be cut in half. He called these particles _____ from the Greek word "atomos", which means _____.

4. John _____ said that atoms were the smallest part of an element that retains its chemical properties.

5. Atoms are made up of three subatomic particles called _____, _____, and _____.

6. The atom's center or _____ is a cluster of protons and neutrons. Protons have a _____ electrical charge and neutrons have ____ electrical charge. They are held together by a _____ force, which keeps positively charged protons from _____ each other.

7. The nucleus is surrounded by a cloud of electrons moving near the speed of _____. The electrons have a negative charge that _____ them to the protons and they live in _____ representing different energy levels.

8. On the periodic table, atoms are arranged by _____ _____, which is really the number of protons or electrons. You can find the _____ _____ and the element's _____ on the periodic table.

9. Different types of atoms can bond together to make _____. Water is created by _____ atoms of hydrogen and _____ atom of oxygen.

10. If the nucleus of an atom was the size of an apple, the electrons would be _____ football fields away.

Word Bank for #2-10
Atomic Mass
Atomic Number
Atoms
Attracts
Dalton
Democritus
Electrons
Elements
Five
Indivisible
Light
Molecules
Neutrons
No
Nucleus
One
Periodic Table
Positive
Protons
Repelling
Shells
Strong
Symbol
Two

Try the quiz to complete each statement.

11. The _____ is the smallest unit or part of an element.
12. Scientists have discovered or created over _____ elements.
13. A _____ is two or more atoms bonded together.
14. What is the most abundant atom in the universe? _____

T. Trimpe 2002

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