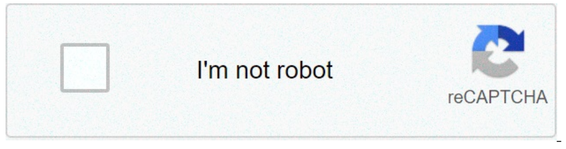

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50 formula of compounds by criss cross method

What is the crossover rule? How do you get the number of oxidation? How do you write chemical formulas? What do you mean by Valence? How do you get equations? What is the Polyatomica Ion in Chemistry? How do you exchange and release in chemistry? Who invented the Criss Cross method? Review of the main arguments Chapter 4 Chemical formulas These skills are usually tested on the subject of SAT in chemistry. You should be able to. . . Ã ¢ â,¬ â ¢ Call up and use the basic rules of oxidation states to write correct formulas. This includes writing formulas with polyatomic ions. Ã ¢ â,¬ â ¢ Compounds of the name (acids, bases and salts) using the equity system and the prefix system and write their formulas. Ã ¢ â,¬ â ¢ Calculate the oxidation states of an element in any formula. Ã ¢ â,¬ â ¢ Calculate the formula mass of a mixture and the percentage composition of each element. Ã ¢ â,¬ â ¢ Calculates the empirical formula When the percentage composition of each element is given. When the mass of the formula is given, you should be able to find the true formula. Ã ¢ â,¬ â ¢ Write a simple equilibrated equation, indicating the phase (or the status) of reagents and products. This chapter will examine and strengthen these skills. Make sure you do practice exercises at the end of the chapter. With the knowledge of the atomic structure, the meaning of the positioning of each element in the periodic table and the link of atoms in ionic and covalent agreements, now you can use this information to write appropriate formulas and appoint the resulting products. Obviously, many compounds can be. A system of writing of the names and formulas of these many combinations has been necessary. The system explained in this text is an organized way to accomplish this. It uses three categories: categories I â,¬ "Binary ionic compounds in which the metal present constitutes only a single type of category II II binary ion compounds with positive charge in which the metal forms more than a type of ion compound with a given In a negative way the charged ions (anion) category III - The covalent binary compounds formed between two non-metallic table 6 is a list of ions that often meet in a chemistry course of the first year. You should know them. Although the use of the Table of the period can help you write the symbol and the apparent charge of attacks and anions, knowing that these common ions can help you write formulas and equations. Category I â,¬ "Binary Ionic Compounds Category The binary ionic compounds are metal ions of groups 1 and 2 of the periodic table. These metal ions have only one type of charge. The formed binary ionic compounds are composed of a positive ion (cation) that is written before and a negative ion (anion). The following rules show how to appoint and write formulas for binary ion compounds. Cacl2 is used as an example. 1. First indicates the action and then anion. 2. Monoatomic action (One-Atom) takes its name from the name of the element. Therefore the soccer ion, the CA2 +, is called football and its chemical symbol comes first. 3. The monoatomical anion with which the action combines is appointed by taking the root of the name of the element and adding Ã ¢ â,¬ "IDE. You must know this rule. The name of the anion is the second. Therefore, l. 'Ionic of chlorine, clÃ ¢', it is called chloride. 4. The name of this mixture is calcium chloride. A quick way to determine the formula of a binary ion compound is to use the Criscross rule. Example 1: To determine the formula for calcium chloride, first write ionic shapes with their associated accusations. Then move the numeric value of the Ion Ion superscript (without fee) to the Pedice of the non-metallic symbol. Then take the numeric value of the SuperScript not. Metallic and make it the metal pedice as shown above. Note that the Numeric 1 is not shown in the final formula. Now you have chlorine 1 as the Pedice of football and football 2 as the chloride's Pedice. As a result, you have CACL2 as a final formula for calcium chloride. Example 2: Write the name name Formula for the formed product when aluminum reacts with oxygen. First write the name. 1. First indicates the action and then anion. 2. Monoatomic action (One-Atom) takes its name from the name of the element. Therefore the aluminum ion, AL3 +, is called aluminum and its chemical symbol appears first. 3. The monoatomical anion with which the action combines is named by taking the root of the element name and adding Ã ¢ â,¬ "IDE. You need to know this rule. The name of the anion is the second. So the ion oxygen, O2Ãâ, ', it is called oxide. 4. The name of this mixture is the aluminum oxide. To determine the formula for aluminum oxide, first write ionic shapes with their associated accusations. Next Move the numerical value of the SuperScript of All '(without charging) to the Pedice of the O. Make the same with the 2 O. in other words, cross the values. Now you have the 2 as an index of aluminum and 3 as an index oxygen. Now you have al2o3 as a final formula for aluminum oxide. This crissoss rule generally works very well. In a situation, but you have to be careful. Suppose you want to write the compound size when magnesium reacts with the Oxygen. Magnesium, an alkaline metal metal in group 2 forms a cation of 2+ and oxygen forms an oion 2. would provide its formula to be mg2o2, but this is not correct. After making the crosscrossing (unless you know that the compound actually exists, as H2O2), it is necessary to reduce all the subscribers from a common factor. In this example, it is possible to divide all subscribers for 2 to obtain the correct formula for magnesium oxide, MGO. When you try to write a formula, you should know if the substance actually exists. For example, you can easily write the formula for carbon nitrate, but no chemist has ever prepared this compound. Tip Remember to reduce all users of a common factor unless you are sure that the mixture exists, like H2O2. _____ * Ionian expenses are shown as numerical exponents followed by the charge. Category II - binary ionic compounds of category II in binary ionic compounds in category II, metals form more than one ion, each with a different charge. The metallic ions (cation) tie IIonically with an ion with charge negatively (anion). The following graph lists most metals that form more than a type of ion cation and therefore more than a binary ion compound with a given anion. _____ * This form of mercury ion (i) occurs always glued as an HG22 + ion. Although the following metals are Ã ¢ â,¬ Â "transitionÃ ¢ â,¬ metals, they form only a type of cation. So a Roman number is not used when naming their compounds. AG1 + Ã, â,¬ æ'silver cd2 + Ã, æ'cadmium zn2 + Ã, â,¬ _ example example: the compound containing the fe2 ion + and the compound containing the fe3 ion + combine with the ionic Chloride to form two different compounds. Using the Crisscross system, get the FECL2 formula for iron chloride (II). The compound formed using the fe3 + ion and the chloride ionic is FECL3, which is the iron chloride (iii). Iron names (ii) chloride and iron (iii) chloride arrived â €

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